

LECTURE NOTE

by Prof. Kook Joe Shin
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Textbook: D. W. Oxtoby, H. P. Gillis, and A. Campion,
"Principles of Modern Chemistry", 7th ed., Cengage, (2012)

Chapter 05 Quantum Mechanics and Atomic Structure

1

Hydrogen Atom

Energy

Levels

- Schrödinger equation for the hydrogen atom
(or one-electron atoms) exactly solvable !

$$-\frac{\hbar^2}{2\mu} \nabla^2 \psi(r) - \frac{Ze^2}{4\pi\epsilon_0 r} \psi(r) = E\psi(r)$$

Reduced mass: $\mu =$

$$\frac{m_1 m_2}{m_1 + m_2}$$

Energy
values

$$E_n = -\frac{Z^2 e^4 m_e}{2n^2 \hbar^2}$$

$$Z^2$$

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