

# CHEMISTRY

## formulas

### TEMPERATURE CONVERSIONS

$$F^{\circ} \rightarrow C^{\circ}$$

$$C^{\circ} = \frac{5}{9}(F^{\circ} - 32)$$

$$C^{\circ} \rightarrow F^{\circ}$$

$$F^{\circ} = \frac{9}{5}C^{\circ} + 32$$

$$C^{\circ} \rightarrow K^{\circ}$$

$$K^{\circ} = C^{\circ} + 273.15$$

$$K^{\circ} \rightarrow C^{\circ}$$

$$C^{\circ} = K^{\circ} - 273.15$$

### AVOGADRO'S NUMBER

$$1 \text{ mole} = 6.02 \times 10^{23}$$

### PERCENT COMPOSITION

$$\text{percent composition} = \frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}}$$

### CHARLE'S LAW

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

V = volume  
T = temperature

### COMBINED GAS LAW

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

P = pressure  
V = volume  
T = temperature

### DILUTION

$$M_1 V_1 = M_2 V_2$$

M = molarity  
V = volume

### PERCENT ERROR

$$\frac{\text{Actual value} - \text{Measured value}}{\text{Actual value}} \times 100\%$$

### MOLE $\rightarrow$ MASS

$$\text{mole} \times \text{mm} = \text{grams}$$

### MASS $\rightarrow$ MOLE

$$\frac{\text{mass}}{\text{mm}} = \text{mole}$$

### TITRATION

$$N_a V_a m_a = N_b V_b m_b$$

V = volume

M = molarity

N = moles

### GAY-LUSSAC'S LAW

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

P = pressure  
T = temperature

### Boyle's Law

$$P_1 V_1 = P_2 V_2$$

P = pressure  
V = volume

# Basic Chemistry

**Judd E. Hollander**



## Basic Chemistry:

**Introduction to Basic Chemistry** Steven S., Zumdahl, Donald J. DeCoste, 2010-02-01 The Seventh Edition of INTRODUCTION TO BASIC CHEMISTRY 7e International Edition includes the first 16 chapters of INTRODUCTION TO CHEMISTRY A FOUNDATION 7e International Edition This new edition features a more explicit problem solving pedagogy to help students develop critical thinking tools they can use in the introductory course and beyond Capturing student interest through early coverage of chemical reactions accessible explanations and visualizations and an emphasis on everyday applications the authors explain chemical concepts by starting with the basics using symbols or diagrams and concluding by encouraging students to test their own understanding of the solution This step by step approach has already helped hundreds of thousands of students master chemical concepts and develop problem solving skills The book is known for its focus on conceptual learning and for the way it motivates students by connecting chemical principles to real life experiences in chapter opening discussions and Chemistry in Focus boxes The Seventh Edition now adds a questioning pedagogy to in text examples to help students learn what questions they should be asking themselves while solving problems offers a revamped art program to better serve visual learners and includes a significant number of revised end of chapter questions The book's unsurpassed teaching and learning resources include a robust technology package that is now supported by OWL Online Web Learning

**Understanding Basic Chemistry: The Learner's Approach** Kim Seng Chan, Jeanne Tan, 2014-08-13 This book the fourth in a series of Understanding Chemistry books deals with Basic Chemistry Written for students taking either the University of Cambridge O level examinations or the GCSE examinations this textbook covers essential topics under both stipulated chemistry syllabi The book is written in such a way as to guide the reader through the understanding and applications of basic essential chemical concepts by introducing a discourse feature the asking and answering of questions that stimulates coherent thinking and hence elucidates ideas Based on the Socratic Method questions are implanted throughout the book to help facilitate the reader's development in forming logical conclusions of concepts The book helps students to master fundamental chemical concepts in a simple way

**An Introduction to General Chemistry** William Martin Blanchard, 1928 **General Chemistry for Colleges** B. Smith Hopkins, 1930 Lecture Notes on General Chemistry Elijah Paddock Harris, 1888 General Chemistry Horace Grove Deming, 1925 *Introduction to General Chemistry* Herbert Newby McCoy, Ethel Mary Terry, 1920 Basic Chemistry of Life Milton Toporek, 1968 *Notes on a Course of Lectures on General Chemistry, Electro-therapeutics and Chemical Physics Delivered by J.W. Langley ...* John Williams Langley, 1882 *General Chemistry for Colleges* Alexander Smith, 1908 **Outline of Lecture Notes on General Chemistry** John Tappan Stoddard, 1885 **General Chemistry** James Kendall, Alexander Smith, 1927 *General Chemistry* Hamilton Perkins Cady, 1916 *A Course of Experiments in General Chemistry and Introduction to Chemical Analysis* Jakob Volhard, C. Zimmermann, 1889 **A Laboratory Outline of General Chemistry** Alexander Smith, 1899

Outlines of General Chemistry Wilhelm Ostwald, 1890      Elements of General Chemistry John Harper Long, 1898  
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