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**G Psacharopoulos**



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monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until complete  
neutralization is obtained

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to  
form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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determine the concentrations of solutions of acids and bases through the titration process we are able to identify physical  
changes to the mixture such as the colour change to indicate the end point of the experiment

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base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known  
molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this  
was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the

end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10.11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid a strong acid with sodium hydroxide a strong base and acetic acid a

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1  $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$  one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $H_2C_2O_4 \cdot 2H_2O$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak  $pK_a$  of an unknown acid or  $pK_b$  of the unknown base key terms titration a process where a

solution of known strength is added to a certain volume of a treated sample containing an  
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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh  
reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4 \text{ aq} + 2\text{NaOH} \text{ aq} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 4\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a ph scale is  
used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally  
ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  
 $\text{NaOH} \text{ aq} + \text{C}_6\text{H}_4\text{COOH} \text{ aq} \rightarrow \text{C}_6\text{H}_4\text{COONa} \text{ aq} + \text{H}_2\text{O} \text{ l}$  the point of equivalence is the point of titration when the number of  
moles of naoh na added is equal to the number of moles of khp nb in the