
Hesi A2 Chemistry Exam Guide 2024

Graded A

Mixture with uniform density throughout and no distinguishable components. - Homogeneous Mixture

Mixture in which the components are readily distinguished. - Heterogeneous Mixture

A change in which the chemical composition of a substance remains the same. - Physical Change

A change in which the chemical bonds are broken and reformed to create a new and different substance. - Chemical Change

The simplest of substances and is represented by a specific letter or combinations of letters. - Element

Combinations of elements in whole number ratios. - Compounds

Mass cannot be created or destroyed during a chemical reaction. - Law of Conservation of Mass

The breaking of bonds and the reforming of new bonds to create new chemical compounds with different chemical formulas and different chemical properties. - Chemical Reaction

synthesis, decomposition, combustion, single replacement, double replacement - Five main types of chemical reactions.

Two elements combine to form a product. - Synthesis Reaction

The breaking of a compound into component parts. - Decomposition Reaction

The reaction of a compound or element with oxygen. In the combustion of a hydrocarbon carbon dioxide and water are produced. - Combustion Reaction

Reaction involving ionic compounds. The reactivity of the ionic compounds determines whether the reaction will take place or not. Can be single replacement or double replacement reaction. - Replacement Reactions

Reaction between a more active metal reacting with an ionic compound containing a less active metal to produce a new compound ex. copper wire reacting with aqueous silver nitrate. - Single Replacement Reaction

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Rob Ritchie



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