

ATOMIC PHYSICS: ATOMIC STRUCTURE

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1. HISTORICALLY FAMOUS MODELS OF AN ATOM

John Dalton believed that matter is composed of very small, indivisible things, he called these things "Atoms". It wasn't a completely alien concept, ancient Greeks had held onto the same believe that matter consisted of tiny objects. In those days no body knew of electrons, protons or neutrons and therefore nobody had any idea, factual or intuitive, about the internal structure of an atom.

After J.J. Thomson discovered the electron in 1897, it was realized that atoms might have an internal structure. Which meant that they are not fundamental unit upon which matter is built. However it was unknown how the electrons were distributed in the atom. Proton or the nucleus hadn't been discovered yet but people knew that since atom was electrically neutral over all, because most objects in daily life were electrically neutral, there had to be equal number of positive and negative charges. In 1904 the, so-called, "Plum Pudding Model" was put forth for the internal structure of the atom. In this model the negative electrons were assumed to float in a soup of positive charge, much like plums in a pudding or raisins in a fruit cake.

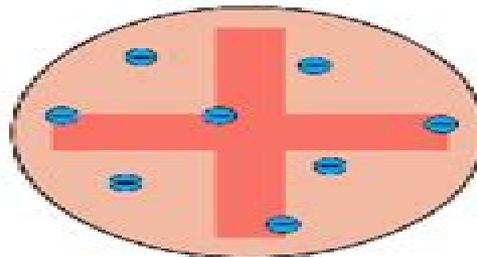


FIGURE 1. The atom according to the Plum Pudding model. Image credit: Wikipedia

In 1896 Henri Becquerel accidentally discovered radioactivity (he intended to discover something else entirely). The same phenomena was then studied by Marie

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